

EXP. NUMBER 5	EXPERIMENT/SUBJECT Kinetic Analysis of Tyrosinase	DATE 7/8+15/09	24
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Title: Kinetic Analysis of Tyrosinase Enzyme  
Reference: Experiment #5 in Modern Exp. Biochem 3<sup>rd</sup> Edition  
 Slides on blackboard.

Synopsis: To examine enzyme kinetics using Tyrosinase enzyme's reaction with Dopa substrate. Ideal concentrations were first determined by measuring the reactions rate via 475nm absorbance (dopa chrome, the product, ideal absorbance). Once determined, the ideal enzyme & substrate concentrations were reacted with an inhibitor (~~the~~ cinnamic acid). The results were used to compare the stereospecific preference of Tyrosinase <sup>mushroom</sup> to either L or D Dopa, and to determine the degree and type of inhibition cinnamic acid is. to determine kinetic constants  $K_m$ ,  $V_{max}$ , &  $K_D$  observations

Procedure

Tyrosinase Concentration was predetermined by Bush using UV 282nm absorbance

$$\frac{1.0\%}{24.9} = \frac{x\%}{0.008} \cdot 3.213 \times 10^{-4} \% \text{ g/100mL}$$

5 assays were set up w/ varying levels of Tyrosinase, a steady level of 3.0 mL total was maintained using phosphate buffer. Absorbance @ 475nm was noted every 30 seconds for 5 min, as well as a 1.

assay #	1	2	3	4	x
buffer (mL)	1.45	1.3	1.1	1.0	
L-Dopa (mL)	1.5	1.5	1.5	1.5	
enzyme (mL)	.05	.2	.4	.5	

\* - Note, only 4 assays done to save enzyme

the buffer + substrate were added first to validate no reaction took place for a minute, then enzyme added.

Slope  $\frac{0.385 - 0.082}{.1}$

Slope  $\frac{0.102}{.1}, \frac{0.385}{.1}, \frac{0.909}{.1}, \frac{1.053}{.1}$

the slopes were analysed looking for a  $\frac{\Delta A}{\Delta t} > .10$

See data table for full results,

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### Procedure

Once ideal enzyme amount assay determined, ideal substrate range was analyzed, using 0.5 mL of enzyme. Validating linearity by looking at  $R^2$  of graph.

### Observations

	1	2	3*
buffer (mL)	2.4	1.5	1.0
L-Dopa (mL)	.1	1.0	1.5
tyrosinase (mL)	.5	0.5	0.5
Slope	.054	.121	.105
$R^2$	<del>.993</del>	.999	.9973

\* - used from enzyme conc determination.

Using the ideal enzyme concentration (0.5 mL) + a good range of substrate (1 to 1.5 mL) measure uninhibited & inhibited rates.

uninhibited

See data tables

inhibited

Validate the inhibitor amount reduces rate by at least 20%

Calculations - Note all performed in excel. sample calculations given + excel printout provided

Substrate concentration  $[S] = \frac{(\text{substrate mL in assay})(\text{stock conc mg/mL})}{3.00 \text{ mL}} \left( \frac{1 \text{ mol Dopa}}{197.23} \right)$

Sample calc. assay #1, uninhibited  $[S] = \left( \frac{(0.1 \text{ mL})(1.515 \text{ mg/mL})}{3.00 \text{ mL}} \right) \left( \frac{1 \text{ mol}}{197.23} \right) = 2.56 \times 10^{-4} \text{ mol/L}$

Rxn Rate,  $V_0$ , calculation

$$V_0 = \left( \frac{(\Delta A / \text{min})}{(3600 \text{ L/mol} \cdot \text{cm})(1.0 \text{ cm})} \right) \left( \overset{\text{volume}}{0.003 \text{ L}} \right) \left( \frac{10^6 \text{ } \mu\text{mol}}{1 \text{ mol}} \right) = \frac{\mu\text{mol}}{\text{min}}$$

Sample calc. assay #1, uninhibited  $V_0 = \left( \frac{.087}{(3600)(1.0)} \right) (0.003) (10^6) = 0.725 \frac{\mu\text{mol}}{\text{min}}$

Calculations (cont)

$V_{max}$   $K_m$ , based on uninhibited plot

~~$K_m$~~   
 $V_{max} = \frac{1}{y_{int}} = \frac{1}{0.7529} = 1.328$   
 Constant units?  $\left(\frac{1}{\mu}\right)$   
 $V_{max} = 2.656 \times 10^{-4}$

Concentration of inhibitor  
 $[I] = \frac{\text{Inhibitor mass}}{\text{Volume}} = \frac{(0.5 \text{ mL})(3.37 \times 10^{-3} \text{ M})}{3.00 \text{ mL}} = 5.62 \times 10^{-4} \text{ M}$

$K_i$ , non competitive  
 2 ways to calc.

$K_i \cdot \text{slope}_{inh} = \text{slope}_{uninh} \left(1 + \frac{[I]}{K_i}\right) = 1.0005 = 1.0002 \left(1 + \frac{5.62 \times 10^{-4}}{K_i}\right) = 3.747 \times 10^{-4} \text{ M}$

OR  
 $(y_{int})_{inh} = (y_{int})_{uninh} \left[1 + \frac{[I]}{K_i}\right] = \frac{1.2405}{0.7529} = 1.2405 \left(1 + \frac{5.62 \times 10^{-4}}{K_i}\right) = 8.678 \times 10^{-4} \text{ M}$   
 average  $K_i = \frac{(3.747 \times 10^{-4}) + (8.678 \times 10^{-4})}{2} = 6.212 \times 10^{-4} \text{ M}$

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